

Fundamentals of Chemistry

Important MCQs



1. There are-----known elements.

- A. 101
- B. 109
- C. 118
- D. 126

ANSWER.C

2. Atoms or groups of atoms with unpaired electrons are known

.....

- A.compounds
- B.mixture
- C.molecules
- D.free radicals

ANSWER.D

3. A gas contains -----molecules per mole.

- A. 6.022×10^{23}
- B. 6×10^{24}
- C. 5.3×10^{24}
- D. 9×10^{27}

ANSWER.A

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4. A gas possesses ----- molecules in two moles.

- A. 1.2×10^{24}
- B. 12×10^{23}
- C. Both
- D. None of them

ANSWER. C

5. Rockets use--- gas for fuel.

- A. Chlorine
- B. hydrogen
- C. Helium
- D. Neon

ANSWER.B

6. The primary ingredient in toothpaste is -----

- A. NaOH
- B. NaHCO₃
- C. Na₂CO₃
- D. H₂O

ANSWER. B

7. The simplest ratio of elements in a formula is given by-----

- A. Empirical formula
- B. Molecular formula
- C. Formula Unit

D. Atomic Number

ANSWER. A

8. The number of moles in 60 grams of carbon

A. 5

B. 6

C. 7

D. 10

ANSWER. A

9. The relative atomic mass in grams is -----

A. Mole

B. Formula Unit

C. Candela

D. Tesla

ANSWER. A

10. One mole of water weighs -----

A. 36g

B. 48g

C. 18g

D. 88g

ANSWER.C

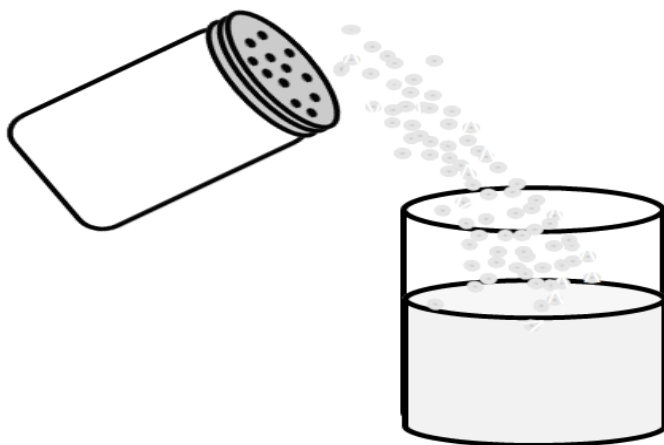
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11. Mixture of salt and water is

- A. Homogeneous
- B. Heterogeneous
- C. Colloidal
- D. Nil



ANSWER. A

12. ----- elements were discovered till 1974

- A.107
- B.105
- C.119
- D.123

ANSWER.B

13. ----- cannot be decomposed into simpler substances

- A. Compound
- B. Solution
- C. Element
- D. Mixture

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ANSWER. C

14. Which of the following is a compound

- A. Carbon
- B. Salt
- C. Argon
- D. Iron

ANSWER.B

15. The valency of oxygen atom is

- A. 6
- B. 2
- C. 4
- D. 16

ANSWER. B

16. Proton number of Carbon is-----

- A.7
- B.6
- C.9
- D.5

ANSWER.B

17. Neutrons carry a ----- charge

- A. +

- B. —
- C. No charge
- D. Double Positive

ANSWER. C

18. Hydrochloric Acid is a ----- molecule

- A. Homoatomic
- B. Heteroatomic
- C. Monoatomic
- D. None of the above

ANSWER. B

19. The relation between atomic number [z] and atomic mass is given by

- A. $A=Z+N$
- B. $A=Z-N$
- C. $A=Z$
- D. $2Z-N$

ANSWER. A

20. One oxygen atom is ---- times heavier than hydrogen

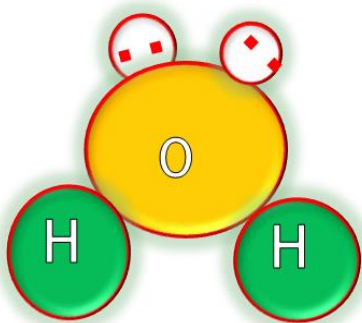
- A. 12
- B. 16
- C. 22
- D. 32

ANSWER. B

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21. Which of the following is a triatomic molecule-----

- A. HCl
- B. H₂O
- C. BF₃
- D. CCl₄



ANSWER. B

22. The unit of relative atomic mass is-----

- A. amu
- B. KJ/mol
- C. Both
- D. None of these

ANSWER. A

23. The amount of matter contained in a body is called its-----

- A. Mass
- B. weight
- C. Molecule
- D. Atom

ANSWER. A

24. The molar mass of NaHCO_3 is-----

- A. 84g
- B. 44g
- C. 40g
- D. 66g

ANSWER.A

25. Molar mass of CaCO_3 is-----

- A. 100g
- B. 96g
- C. 64g
- D. 52g

ANSWER. A

26. Mixtures are of ----- types

- A. 3
- B. 8
- C. 2
- D. 6

ANSWER. C

27. ----- is a charged specie

- A. Ion
- B. Molecule

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- C. Atom
- D. None of the above

ANSWER. A

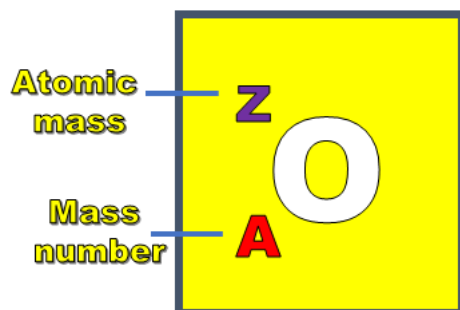
28. Which one is an example of a heterogeneous mixture-----?

- A. Sand in water
- B. Salt in water
- C. Both
- D. None

ANSWER. A

29. Alphabet -----is used to represent the atomic number

- A. C
- B. Z
- C. V
- D. P



ANSWER. B

30. Usually an atom contains -----number of protons and electrons

- A. Equal
- B. Unequal
- C. Cannot be determined
- D. Unknown

ANSWER. A

31. The symbol used to denote tin is-----

- A. Sn
- B. O
- C. F
- D. U

ANSWER. A

32. The percentage of Ca in CaCO₃ is-----

- A. 35%
- B. 40%
- C. 55%
- D. 67%

ANSWER. B

33. The number of neutrons can be determined by-----

- A. X-Z
- B. A-Z
- C. V-B
- D. X-A

ANSWER. B

34. The empirical formula of glucose is-----

- A. CH₂O
- B. C₆H₆
- C. H₂O
- D. V₂O₅

ANSWER. A

35. The empirical formula of hydrogen peroxide is-----

- A. H₂O₂
- B. KO₂
- C. TI₂O₅
- D. HO

ANSWER. D

36. Benzene is a ----- molecule

- A. Heteroatomic
- B. Homoatomic
- C. Both
- D. None

ANSWER. A

37. Number of naturally occurring elements is-----

- A. 87
- B. 99

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C. 92

D. 55

ANSWER. C

38. The smallest unit of an ionic compound is -----

A. Ion

B. Electron

C. Both

D. None

ANSWER. A

39. Which one of the following discovered Avogadro number

A. Marie Curie

B. Amadeo Avogadro

C. Tesla

D. Charles Darwin

ANSWER. B

40. Which of the following are more electronegative-----

A. Metals

B. Non-metals

C. Metalloids

D. None

ANSWER. B

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41. Noble gases have the most ----- electronic configuration

- A. Unstable
- B. Stable
- C. Both
- D. None

ANSWER. B

42. Which of the following is the most abundant gas in the atmosphere -----

- A. Nitrogen
- B. Oxygen
- C. Helium
- D. Hydrogen

ANSWER. A

43. Percentage of nitrogen gas in the atmosphere is-----

- A. 78%
- B. 56%
- C. 44%
- D. 87%

ANSWER. A

44. Mass of one mole of NaCl is-----

- A. 55g
- B. 57.5g
- C. 87g
- D. 43g

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ANSWER. B

45. The number of moles in 54g of water is-----

- A. 5
- B. 3
- C. 7
- D. 9

ANSWER. B

46. The chemical formula of methane is

- A. CH₂
- B. CH₃
- C. CH₄
- D. C₂H₆

ANSWER. C

47. The percentage composition of C in CH₄ is

- A. 67%
- B. 88%
- C. 75%
- D. 34%

ANSWER. C

48. One mole of oxygen atom is-----

- A. 32g
- B. 18g

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C. 16g

D. 19g

ANSWER. C

49. Mass of 1 mole of Benzene is----

A. 88g

B. 78g

C. 98g

D. 54g

ANSWER. B

50. 92g of nitrogen dioxide contains -----mole

A. 4

B. 2

C. 7

D. 8

ANSWER. B

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