

Chemistry Important MCQs

Class 9 Chapter 6

Solutions

Multiple Choice Questions.

1. A solution with uniform composition is known as

- A. Saturated
- B. Unsaturated
- C. Homogeneous
- D. Heterogeneous

ANSWER. C. Homogeneous

2. The substance present in relatively lesser amount is known as

- A. Solute
- B. Solvent
- C. Medium
- D. None of these

ANSWER; A. Solute

3. A solution in which ----- is a solvent is known as an aqueous solution.

- A. Hydrochloric acid
- B. Acetic acid

C. Water

D. Ammonia

ANSWER; C. Water

4. The component of solution with a relatively higher concentration is

A. Solvent

B. Solute

C. Medium

D. All of these

ANSWER; A. Solvent

5. The word aqueous is derived from a ----- word.

A. Spanish

B. French

C. Latin

D. Dutch

ANSWER; C. Latin

6. The solution which can dissolve more solute is known as ----- solution

A. Unsaturated

B. Saturated

C. Super Saturated

D. All of These

ANSWER; A. Unsaturated

7. The solution which cannot dissolve more solute is known as ----- solution

A. Unsaturated

B. Saturated

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C. Super Saturated

D. All of These

ANSWER; B. Saturated

8. A solution that contains more solute than a saturated solution is known as ----- solution.

A. Unsaturated

B. Saturated

C. Super Saturated

D. All of These

ANSWER; C. Super Saturated

9. Sodium chloride in water is a prime example of ----- solution.

A. Aqueous

B. Acidic

C. Basic

D. None of these

ANSWER; A. Aqueous

10. Sand in water is an example of ----- solution.

A. Saturated

B. Unsaturated

C. Homogeneous

D. Heterogeneous

ANSWER; D. Heterogeneous

11. Super Saturated solution is formed by ----- a saturated solution.

- A. Diluting
- B. Heating**
- C. Mixing
- D. None of these

ANSWER; B. Heating

12. The ----- that makes up the air we breathe includes rare gases, nitrogen, oxygen, and carbon dioxide.

- A. Solution
- B. Gaseous mixture**
- C. Colloids
- D. All of these

ANSWER; B. Gaseous mixture

13. Fertilizer is a gaseous mixture of NH_3 and -----

- A. HCO_3
- B. NO_2
- C. CO_2**
- D. HNO_3

ANSWER; C. CO_2

14. Fog is a mixture of ----- in air.

- A. Water vapours**
- B. Ammonia
- C. HCl
- D. All of these

ANSWER; A. Water vapours

15. The metal ----- is the only metal in a liquid state.

- A. Tungsten**
- B. Radium**
- C. Copper**
- D. Mercury**

ANSWER; D. Mercury

16. Vinegar is a ----- solution of acetic acid in water.

- A. 10%**
- B. 5%**
- C. 8%**
- D. 5%**

ANSWER; B. 5%

17. The quantity of ----- in a solution is known as concentration.

- A. Solvent**
- B. Medium**
- C. Solute**
- D. None of these**

ANSWER; C. Solute

18. Concentrated solution of HCl contains ----- HCl in 100g of solution.

- A. 98g**
- B. 88g**
- C. 49g**
- D. 37g**

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ANSWER; D. 37g

19. ----- is defined as the number of moles of solute dissolved per dm³ of solution.

A. Molarity

B. Molality

C. Concentration

D. All of these

ANSWER; A. Molarity

20. Potassium chlorate is a ----- coloured solid.

A. Green

B. Orange

C. White

D. Purple

ANSWER; C. White

21. The amount of solute that dissolves 100g of a solvent is known as-----

A. Concentration

B. Molality

C. Molarity

D. Solubility

ANSWER; D. Solubility

22. Solubility is dependent upon -----

A. Temperature

B. Acid

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- C. Water
- D. None of these

ANSWER; A. Temperature

23. Methanol dissolves in water due to -----

- A. Covalent bond
- B. Ionic bond
- C. Hydrogen bonding**
- D. All of these

ANSWER; C. Hydrogen bonding

24. Glucose is ----- soluble in water.

- A. Readily**
- B. Not
- C. Partially
- D. All of these

ANSWER; A. Readily

25. Solubility of ionic compounds increase with the increase in -----

- A. Concentration
- B. Boiling point
- C. Melting point
- D. Temperature**

ANSWER; D. Temperature

26. Solubility of Na_2SO_4 ----- with increase in temperature.

- A. Increases**

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B. Decreases

C. Not predictable

D. All of these

ANSWER; B. Decreases

27. In a homogeneous mixture, the size of particles is between ----- nm.

A. 0.1-1

B. 3-5

C. 7-9

D. 2-3

ANSWER; A. 0.1-1

28. In colloids the particle size is ----- than in heterogeneous mixture.

A. Bigger

B. Not predictable

C. Smaller

D. All of these

ANSWER; C. Smaller

29. A mixture in which solute particles can be easily seen is known as

A. Homogeneous

B. Heterogeneous

C. Colloidal

D. Suspension

ANSWER; D. Suspension

30. The particles of a suspension are easily

A. Homogenized

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- B. Not Visible
- C. Visible
- D. Both A and B

ANSWER; C. Visible

31. Sodium Hydroxide is used in the preparation of

- A. Soaps
- B. Tyres
- C. Pencils
- D. All of These

ANSWER; A. Soaps

32. 0.25M NaOH contains a mass of ----- g.

- A. 50
- B. 43
- C. 77
- D. 10

ANSWER; D. 10

33. KClO₃ is a prime component of

- A. Nail polishes
- B. Soil
- C. Colours
- D. Dyes

ANSWER; D. Dyes

34. Potassium Permanganate (KMnO₄) is soluble in water solution and is ----- color.

- A. Red
- B. Orange
- C. Purple**
- D. Dark blue-black

ANSWER; C. Purple

35. Urea is a starting material for

- A. Plastic
- B. Fertilizer
- C. Both**
- D. None of these

ANSWER; C. Both

36. Pure----- is very soft.

- A. Silver
- B. Platinum
- C. Mercury
- D. Gold**

ANSWER; D. Gold

37. Brine is a mixture of ----- in water.

- A. Glucose
- B. Sucrose
- C. Lactose
- D. Salt**

ANSWER; D. Salt

38. Ozone filters all the ----- radiations.

- A. Infrared
- B. Visible
- C. Ultraviolet**
- D. All of these

ANSWER; C. Ultraviolet

39. A super-saturated solution is ----- in presence of crystals.

- A. Stable
- B. Not predictable
- C. Not stable**
- D. All of these

ANSWER; C. Not Stable

40. German silver is an -----

- A. Solution
- B. Suspension
- C. Alloy**
- D. All of these

ANSWER; C. Alloy
